#### **Chem 111**

### 9:05a section

## Evening Exam #2v1

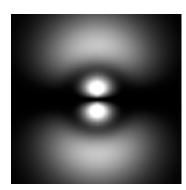
This exam is composed of **25** questions. Go initially through the exam and answer the questions you can answer *quickly*. Then go back and try the ones that are more challenging to you and/or that require calculations.

As discussed on the course syllabus, honesty and integrity are absolute essentials for this class. In fairness to others, dishonest behavior will be dealt with to the full extent of University regulations.

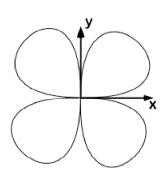
$$E = hv = \frac{hc}{\lambda}$$
  $h = 6.626x10^{-34} J s$   
 $c = 2.998x10^8 m s^{-1}$   
 $N = 6.022x10^{23} mol^{-1}$   
 $N = 6.022x10^{23} mol^{-1}$ 

- 1. How many valence electrons are in the O atom?
  - 1)4
- 2) 6
- 3)8
- 4) 16
- 5) 0
- 2. Which atom(s) has/have completely filled 3s, 3p, and 3d orbitals?
  - 1) Ar
- 2) Zn
- 3) Kr
- 4) Ar & Zn
- 5) Kr & Zn
- 3. Which element is represented by:  $1s^22s^22p^63s^23p^63d^{10}4s^24p^64d^{10}5s^25p^3$ 
  - 1) Sb
- 2) Te
- 3) Br
- 4) As
- 5) Se

- 4. The orbital depicted at right is:
  - 1) 1s
- 2) 2p
- 3) 3s
- 4) 3p
- 5) 4p



- 5. The orbital depicted at right is:
  - 1)  $p_{xy}$  2)  $d_{xy}$  3)  $d_{x}2_{-y}2$  4)  $d_{z}2$  5)  $f_{xy}$



- 6. Which of the following quantum number sets is *not* allowed?

  - 1) n=+3 l=+2  $m_l=-1$   $m_s=+1/2$  2) n=+2 l=+1  $m_l=-1$   $m_s=+1/2$
  - 3) n=+3 l=+1  $m_l=-1$   $m_s=-1/2$  4) n=+2 l=0  $m_l=-1$   $m_s=+1/2$
  - 5) n=+3 l=0  $m_l=0$   $m_s=-1/2$
- 7. What is the maximum number of orbitals that can be identified by the set of quantum numbers n=+3 l=+2?
  - 1) 7
- 2) 2
- 3) 3
- 4) 5
- 5) 10

- 8. Which of the following has the shortest bond length?
  - 1) H<sub>2</sub>O
- 2) HF
- 3) NH<sub>3</sub>
- 4)  $CH_4$
- 9. Which of the following has the highest bond energy?
  - 1) B<sub>2</sub>

- 2)  $C_2$  3)  $N_2$  4)  $O_2$
- 5)  $F_2$

- 10. The CO bond in the molecule CH<sub>3</sub>OH is best described as a:
  - 1) triple bond

2) double bond

3) single bond

- 4) ionic bond
- 5) the molecule doesn't exist

- 11. Consider the molecule  $SO_3^x$ , where x is the charge on the molecule. All three bonds are single bonds. Which value of x yields the stable molecule? (Hint: draw Lewis structures to figure this one out)
  - 1) + 2
- 2) + 1
- 3)0
- 4) -1
- 5)-2

- 12. For the  $SO_3^x$  molecule above, how many equal-energy resonance structures can you draw?
  - 1) 1
- 2) 2
- 3) 3
- 4) 4
- 5)6

- 13. The NO bond in HNO is a:
  - 1) single bond
- 2) double bond
- 3) triple bond
- 4) ionic bond
- 14. If an element with the valence configuration  $4s^23d^7$  loses 2 electron(s), these electron(s) would be removed from the following **subshell(s)**.
  - 1) 4s
- 2) 3d
- 3) 4s and 3d
- 4) 3p
- 5) 4p

15. Which molecule below does not exist?

- 1) BeF<sub>2</sub>
- 2) CaF<sub>4</sub>
- 3) MgO
- 4) KCl

5) BCl<sub>3</sub>

16. Draw a stable Lewis structure for the symmetrical molecule **hydrazine** N<sub>2</sub>H<sub>4</sub>. In this structure, how many *lone pair electrons* are on *each* N?

- 1) 1
- 2) 2
- 3) 3
- 4) 4

5) 6

17. Draw a stable Lewis structure for the molecule **OCS**. In this structure (with C at the center), what is the bond order between C and O?

- 1) 1
- 2) 1.5
- 3) 2
- 4) 2.5

5) 3

18. Draw the best Lewis structure for  $ClF_2^+$ . How many **lone pair electrons** are located on Cl?

- 1) 1
- 2) 2
- 3) 3
- 4) 4

5)6

19. For the molecule  $ClF_2^+$ , what is the electron group geometry of Cl?

1) linear

- 2) tetrahedral
- 3) trigonal planar

- 4) trigonal bipyramidal
- 5) octahedral

20. In the molecule  $NO_2^+$ , the actual bond order for each NO bond is:

1) 1

2) 2

3)3

4) 1.5

5) 1 for one bond and 2 for the other

21. Draw the Lewis structure for  $XeF_4$ . The electron group geometry is:

1) square planar

2) square pyramidal

3) trigonal bipyramidal

4) octahedral

5) none of the above

# Bond Dissociation Energies (kJ mol<sup>-1</sup>) (gas phase)

			_				
Bond	D	Bond	D	Bond	D		
Н-Н	436	C-C	346	N-N	163		
С-Н	413	C=C	610	N=N	418		
N-H	391	O-O	146	C-O	358		
O-H	463	O=O	498	C=O	745		

22. Consider the reaction:  $H_2CCH_2(g) + H_2(g) \Rightarrow CH_3CH_3(g)$ 

What is the energy  $(\Delta H^{\circ}$ , in kJ mol<sup>-1</sup>) for this reaction?

1) –480

2) –44

3) +44

4) -346

5) +346

5) -1 for one O and 0 for the other O

(Questions 23-24) Consider the following resonance forms for the ion OCN

- 23. In resonance structure **b**, what is the formal charge on O?
  - 1) -3
- 2) -2
- 3)-1
- 4) 0
- 5) + 1
- 24. Which resonance structure is higher in energy, **b** or **c**?
  - 1) b
- 2) c
- 3) neither, they have the same energy
- 25. The correct designator for this course is:
  - 1) Chem 111
- 2) Chem 363
- 3) Econ 3.33
- 4) Sports 01

#### PERIODIC TABLE OF THE ELEMENTS

1A	2A	3B	4B	5B	6B	7B	8B	8B	8B	1B	2B	3A	<b>4A</b>	5A	6A	7A	8A
1																	2
H																	He
1.008		_															4.003
3	4											5	6	7	8	9	10
Li	Be											В	C	N	O	F	Ne
6.939	9.012											10.81	12.01	14.01	16.00	19.00	20.18
11	12											13	14	15	16	17	18
Na	Mg											Al	Si	P	S	Cl	Ar
22.99	24.31											26.98	28.09	30.97	32.07	35.45	39.95
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
K	Ca	Sc	Ti	$\mathbf{V}$	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
39.10	40.08	44.96	47.90	50.94	52.00	54.94	55.85	58.93	58.71	63.55	65.39	69.72	72.61	74.92	78.96	79.90	83.80
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
85.47	87.62	88.91	91.22	92.91	95.94	(99)	101.1	102.9	106.4	107.9	112.4	114.8	118.7	121.8	127.6	126.9	131.3
55	56	57	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
Cs	Ba	La	Hf	Ta	$\mathbf{W}$	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
132.9	137.3	138.9	178.5	181.0	183.8	186.2	190.2	192.2	195.1	197.0	200.6	204.4	207.2	209.0	(209)	(210)	(222)
87	88	89	104	105	106	107	108	109									
Fr	Ra	Ac	Unq	Unp	Unh	Uns	Uno	Une									
(223)	226.0	227.0	(261)	(262)	(263)	(262)	(265)	(266)									