

FOR AN “IDEAL” MIXTURE OF TWO LIQUID COMPOUNDS:

AT ANY GIVEN TEMPERATURE, THE VAPOR PRESSURE OF THE LOWER BP COMPOUND > THE VAPOR PRESSURE OF THE HIGHER BP COMPOUND.

THUS, THE VAPOR ABOVE THE LIQUID WILL BE RICHER IN THE LOWER-BOILING COMPOUND, COMPARED TO THE RELATIVE AMOUNTS IN THE LIQUID PHASE.

EXAMPLE: A CYCLOHEXANE/TOLUENE MIXTURE

