



OWL Question

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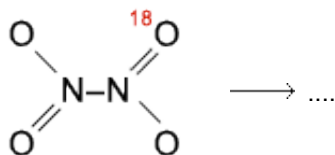
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Chemical Formulas	Scientific Notation	Periodic Table	Tables
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Dinitrogen tetraoxide is a colorless gas that dissociates into nitrogen dioxide, a reddish brown gas.



An experiment was run to demonstrate that this is a dynamic equilibrium. Starting with a special form of N_2O_4 where one of the oxygens was isotopically labeled (^{18}O instead of ^{16}O), the system was then allowed to reach equilibrium.



If the equilibrium is **NOT dynamic** the forward reaction would simply run until equilibrium is reached and then cease. Under this condition which of the structures below would be observed for N_2O_4 ?

If the equilibrium **IS dynamic** the forward and reverse reactions run even after equilibrium is reached. Under this condition which of the structures below would be observed for N_2O_4 ?

Put letters in alphabetical order. If none are correct enter NONE.

<p>A.</p>	<p>C.</p>
<p>B.</p>	<p>D.</p>