

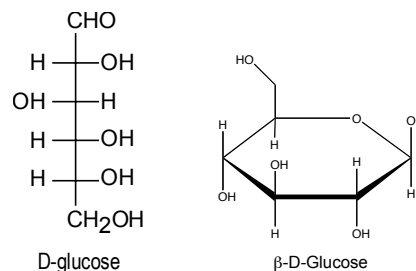
## Chem 250

## Answer Key Evening Exam 2v2

This exam is composed of **40** questions.

As discussed in the course syllabus, honesty and integrity are absolute essentials for this class. In fairness to others, dishonest behavior will be dealt with to the full extent of University regulations.

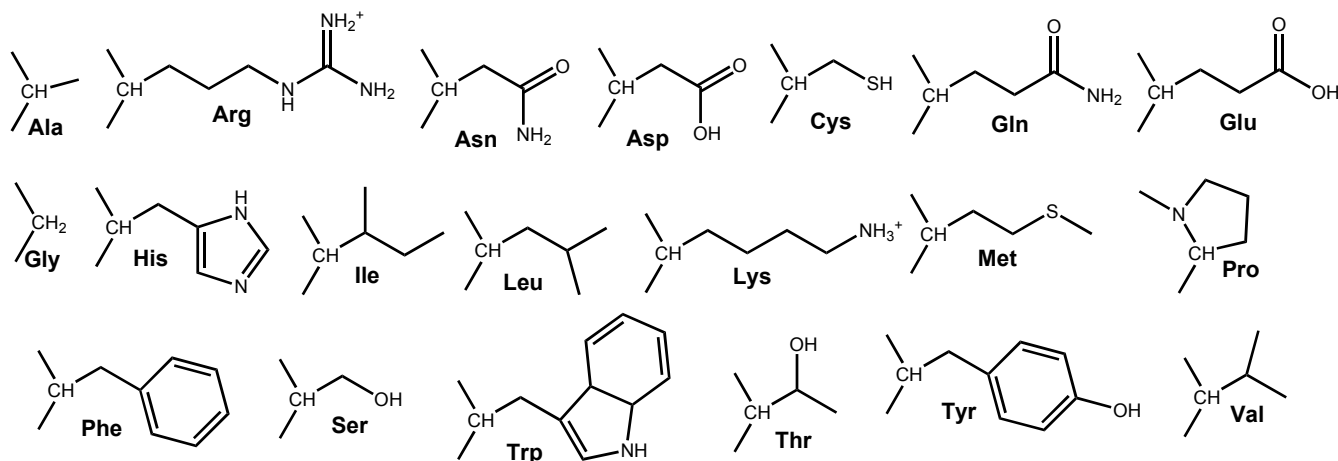
I hereby state that all answers on this exam are my own and that I have neither gained unfairly from others nor have I assisted others in obtaining an unfair advantage on this exam.



\_\_\_\_\_  
Signature

## PERIODIC TABLE OF THE ELEMENTS

1A	2A	3B	4B	5B	6B	7B	8B	8B	8B	1B	2B	3A	4A	5A	6A	7A	8A
1 <b>H</b> 1.008																	2 <b>He</b> 4.003
3 <b>Li</b> 6.939	4 <b>Be</b> 9.012											5 <b>B</b> 10.81	6 <b>C</b> 12.01	7 <b>N</b> 14.01	8 <b>O</b> 16.00	9 <b>F</b> 19.00	10 <b>Ne</b> 20.18
11 <b>Na</b> 22.99	12 <b>Mg</b> 24.31											13 <b>Al</b> 26.98	14 <b>Si</b> 28.09	15 <b>P</b> 30.97	16 <b>S</b> 32.07	17 <b>Cl</b> 35.45	18 <b>Ar</b> 39.95
19 <b>K</b> 39.10	20 <b>Ca</b> 40.08	21 <b>Sc</b> 44.96	22 <b>Ti</b> 47.90	23 <b>V</b> 50.94	24 <b>Cr</b> 52.00	25 <b>Mn</b> 54.94	26 <b>Fe</b> 55.85	27 <b>Co</b> 58.93	28 <b>Ni</b> 58.71	29 <b>Cu</b> 63.55	30 <b>Zn</b> 65.39	31 <b>Ga</b> 69.72	32 <b>Ge</b> 72.61	33 <b>As</b> 74.92	34 <b>Se</b> 78.96	35 <b>Br</b> 79.90	36 <b>Kr</b> 83.80
37 <b>Rb</b> 85.47	38 <b>Sr</b> 87.62	39 <b>Y</b> 88.91	40 <b>Zr</b> 91.22	41 <b>Nb</b> 92.91	42 <b>Mo</b> 95.94	43 <b>Tc</b> (99)	44 <b>Ru</b> 101.1	45 <b>Rh</b> 102.9	46 <b>Pd</b> 106.4	47 <b>Ag</b> 107.9	48 <b>Cd</b> 112.4	49 <b>In</b> 114.8	50 <b>Sn</b> 118.7	51 <b>Sb</b> 121.8	52 <b>Te</b> 127.6	53 <b>I</b> 126.9	54 <b>Xe</b> 131.3
55 <b>Cs</b> 132.9	56 <b>Ba</b> 137.3	57 <b>La</b> 138.9	72 <b>Hf</b> 178.5	73 <b>Ta</b> 181.0	74 <b>W</b> 183.8	75 <b>Re</b> 186.2	76 <b>Os</b> 190.2	77 <b>Ir</b> 192.2	78 <b>Pt</b> 195.1	79 <b>Au</b> 197.0	80 <b>Hg</b> 200.6	81 <b>Tl</b> 204.4	82 <b>Pb</b> 207.2	83 <b>Bi</b> 209.0	84 <b>Po</b> (209)	85 <b>At</b> (210)	86 <b>Rn</b> (222)
87 <b>Fr</b> (223)	88 <b>Ra</b> 226.0	89 <b>Ac</b> 227.0	104 <b>Unq</b> (261)	105 <b>Unp</b> (262)	106 <b>Unh</b> (263)	107 <b>Uns</b> (262)	108 <b>Uno</b> (265)	109 <b>Une</b> (266)									



1. (2 points) Propanoic acid and methyl acetate are constitutional isomers and both are liquids at room temperature. Which has the higher boiling point?

1) propanoic acid      2) methyl acetate      3) they have the same boiling point

**(1) propanoic acid – the carboxylic acid in propanoic acid can form very stable interactions with itself? Methyl acetate is an ester (Chptr 18)**

2. Which listing below correctly orders the boiling points of the indicated molecules?

1) 1-butanol > propanoic acid > diethyl ether

2) propanoic acid > 1-butanol > diethyl ether

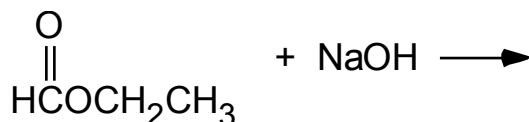
3) propanoic acid > diethyl ether > 1-butanol

4) diethyl ether > propanoic acid > 1-butanol

5) 1-butanol > diethyl ether > propanoic acid

**(2) acid > alcohol > ether (all have about the same MW) (Chptr 18)**

3. (2 points) The products of the following reaction are:



1) sodium propanoate and water

2) sodium acetate and formaldehyde

3) ethanol and sodium formate

4) methane and sodium acetate

5) none of the above

**(3) hydrolysis of an ester Chapter 19, p 480 / Quiz 2**

4. (2 points) Which of the following is expected to have the highest melting point?

1)  $\text{CH}_3(\text{CH}_2)_{20}\text{COOH}$

2)  $\text{CH}_3(\text{CH}_2)_{12}\text{COOH}$

3) it is not possible to predict

**(1) longer chain leads to more van der Waals interactions OWL 18.4a / Quiz 2**

5. (2 points) Glycolipids contain what characteristic head group?

1) sphingosines

2) carbohydrates

3) cholesterol

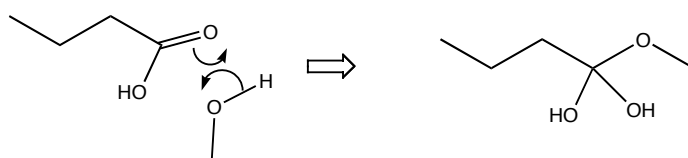
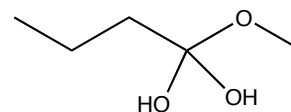
4) steroids

5) phosphates

**(2) Chptr 21**

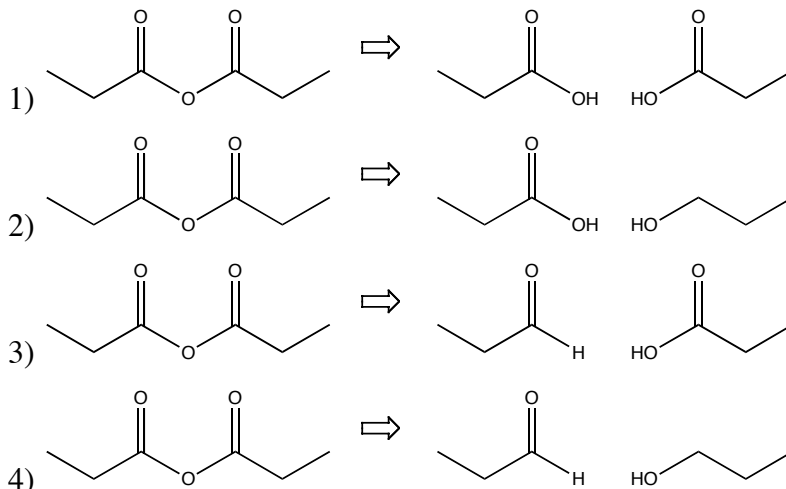
6. (2 points) Which two reactants would lead to the Fischer esterification reaction intermediate shown at right?

- 1) butanal and formic acid
- 2) butanoic acid and methanol
- 3) 1-butanone and formic acid
- 4) pentanoic ester and water
- 5) none of the above



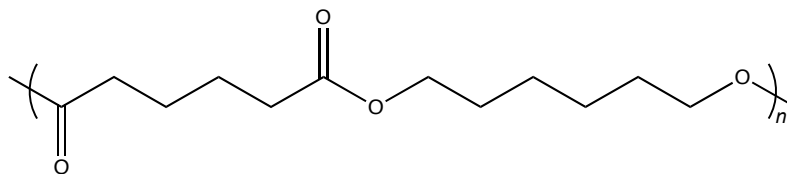
**(2) OWL 18.5e / Quiz 2**

7. (2 points) Hydrolysis of propyl anhydride is represented by which reaction below?



**(1) Chptr 19**

8. (2 points) The molecule below is an example of what kind of polymer?



- 1) polyester
- 2) polycarbonate
- 3) polyamide
- 4) polyacrylate

**(1) (Chptr 19)**

9. (2 points) In one complete cycle of the acyl carrier protein, how many carbons are added to the growing fatty acyl chain?

- 1) 1
- 2) 2
- 3) 3
- 4) 4
- 5) 8

**(2) (Chapter 29)**

10. (2 points) The synthesis of one molecule of glucose requires how many molecules of acetyl-CoA?

- 1) 1                      2) 2                      3) 3                      4) 4                      5) 8

**(3) (Chapter 29)**

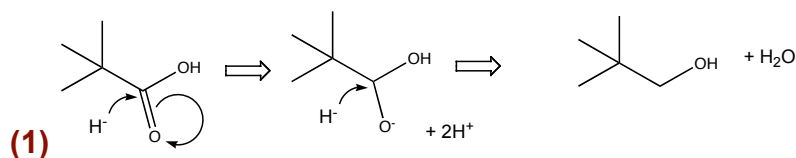
11. (2 points) The reactions of gluconeogenesis are simply the reactions of glycolysis run in reverse

- 1) True                      2) False

**(2) False (OWL 29.2)**

12. (2 points) The reaction of 2,2-dimethylpropanoic acid and  $\text{LiAlH}_4$  in water yields:

- 1) water and 2,2-dimethyl-1-propanol  
 2) water and 2,2-dimethyl-1-propanal  
 3)  $\text{CO}_2$  and 2,2-dimethylbutanoic acid  
 4)  $\text{CO}_2$  and 2,2-dimethylethanoic acid  
 5) nothing. No reaction occurs.



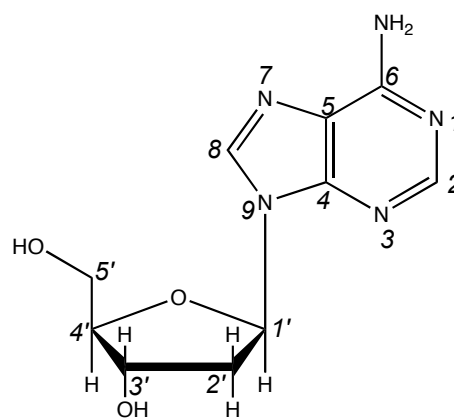
13. (2 points) Which of the following is/are aldose(s)?

- 1)  $\begin{array}{c} \text{CHO} \\ | \\ \text{CHOH} \\ | \\ \text{CH}_2\text{OH} \end{array}$       2)  $\begin{array}{c} \text{CH}_2\text{OH} \\ | \\ \text{CO} \\ | \\ \text{CH}_2\text{OH} \end{array}$       3) Both      4) Neither

**(1) The first one is. see p 494 and OWL 20.1b**  
**Aldose → Aldehyde. Look for the one with the aldehyde (Quiz 2)**

14. (2 points) In adenosine, shown at right, which of the following sugar centers are chiral (note that the sugar atoms are labeled  $n'$ , while the base atoms are labeled  $n$ ).

- 1) 1', 2', 3', 4', and 5'
- 2) 1', 2', 3', 4', and 5'
- 3) 1', 3', 4', and 5'
- 4) 1', 3', and 4'
- 5) 1', 2', 3', and 4'



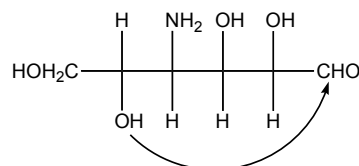
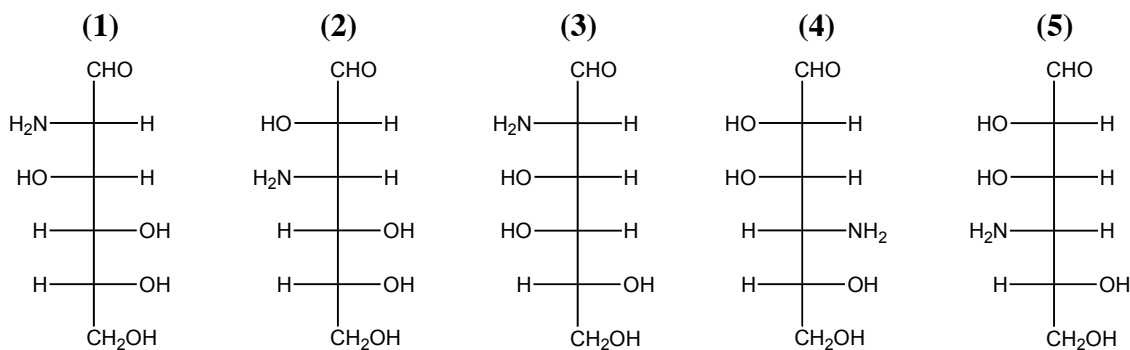
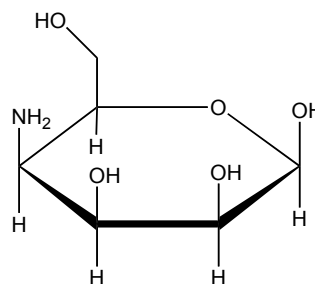
**(4) Chptr 20 and Gen Chem**

15. (2 points) Which of the following atoms in adenosine are  $sp^2$  hybridized?

- 1) 1', 2', 3', 4', and 5'
- 2) 1', 2', 3', and 4'
- 3) 1 through 9
- 4) 1 through 8
- 5) 1 through 9 and 5'

**(3) Chptr 20 and Gen Chem**

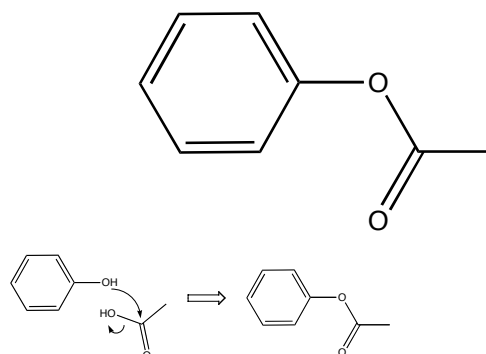
16. (2 points) Which Fischer representation below shows the linear form of the cyclic sugar shown at right?



**(5) Chapter 20 (p 499) / Quiz 2**

17. (2 points) From what parent molecules can the molecule at right be synthesized?

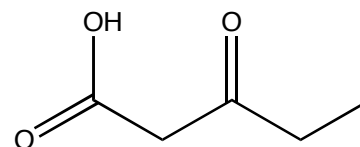
- 1) toluene and methyl acetate
- 2) benzoic acid and methanol
- 3) benzene and acetic acid
- 4) acetic acid and phenol



**(4) - (Chptr 18)**

18. (2 points) Heating the molecule at right yields which products?

- 1) propanoic acid and carbon dioxide
- 2) acetic acid and propanoic acid
- 3) 2-butanone and carbon dioxide
- 4) butanoic anhydride
- 5) no reaction occurs



**(3) decarboxylation of  $\beta$ -ketocarboxylic acid (Chptr 18)**

19. (2 points) In metabolism, CoA-SH usually reacts directly with

- |                     |           |               |
|---------------------|-----------|---------------|
| 1) alcohols         | 2) esters | 3) anhydrides |
| 4) carboxylic acids | 5) water  |               |

**(4) (Chapter 27)**

20. (2 points) In NAD and FAD the adenosine diphosphate functional group serves what purpose?

- 1) Hydrolysis of the diphosphate helps to drive reactions
- 2) It accepts a phosphate from reactants to dephosphorylate them
- 3) It is a “handle” to help it bind to enzyme active sites
- 4) It helps to bind and position the reactants
- 5) It plays no role. You can remove it and what remains still functions

**(3) (Chapter 27)**

21. (2 points) In the Citric Acid cycle, succinate reacts with FAD. In this reaction, succinate:

- 1) isomerizes
- 2) is phosphorylated
- 3) is dephosphorylated
- 4) is oxidized
- 5) is reduced

**(4) (Chapter 27)**

22. (2 points) The negatively charged molecule carbonylcyanide-*p*-trifluoromethoxyphenylhydrazone (FCCP) binds to H<sup>+</sup> ions in the mitochondrial intermembrane space and transports them across the inner membrane to the matrix. FCCP thus is toxic because it:

- 1) prevents electron flow to dioxygen
- 2) leads to the build up of lactic acid
- 3) prevents synthesis of ATP via the proton translocating ATPase
- 4) leads to excess protonation of acetyl-CoA
- 5) inhibits phosphorylation of glucose

**(3) (OWL 27.6)**

23. (2 points) Which listing below contains only hydrophobic amino acids?

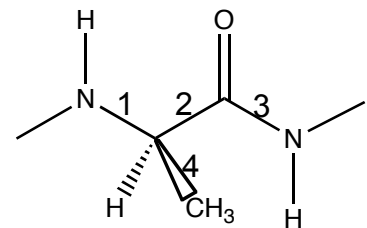
- |                       |                       |
|-----------------------|-----------------------|
| 1) Ile, Leu, Val, Phe | 2) Met, Asn, Pro, Leu |
| 3) Arg, Glu, Asp, Lys | 4) Arg, Glu, Val, Phe |
| 5) Met, Asn, Asp, Lys |                       |

**(1) Chptr 22**

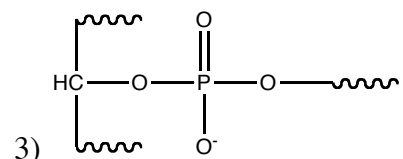
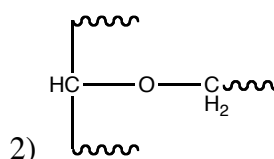
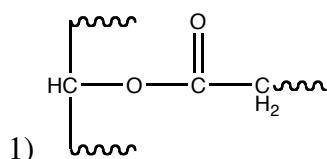
24. (2 points) In the amino acid linkage shown at right, which bonds have a high energy cost for rotation?

- 1) 1                      2) 2                      3) 3                      4) 4

**(3) Remember the resonance structure (Chptr 22)**

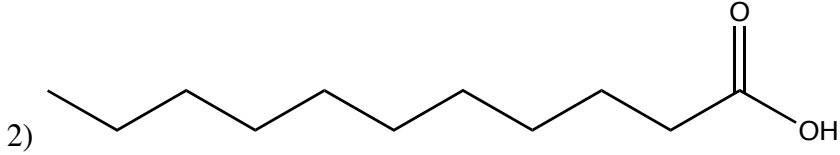
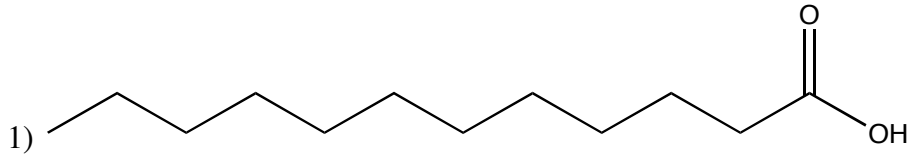


25. (2 points) Which of the following is a structure typically *not* seen in the complex lipids we have seen?



**(2) Chptr 21**

26. (2 points) Which fatty acid below is not of natural origin?



3) Neither are of natural origin

4) Both are of natural origin

**(2) – 11 carbons – even number (Chptr21)**

27. (2 points) Ounce for ounce, which provide higher energy yields, both in conventional combustion and in cellular metabolism

1) carbohydrates

2) fats

**(2) Fats are more reduced and so yield more energy in oxidation to CO<sub>2</sub> (Chapter 28)**

28. (2 points) ATP is often hydrolyzed in order to drive unfavorable reactions. Another important and very common role for ATP that does not involve hydrolysis is:

1) reduction of carboxylic acids

2) oxidation of alcohols

3) phosphorylation of alcohols

4) oxidation of primary amines

5) cyclization of sugars

**(3) Chptr 28**

29. (2 points) What force is most dominant in driving a protein from an ensemble unfolded of states to a compact globular structure?

1) hydrophobic collapse

2) hydrogen bonding

3) disulfide bonding

4) formation of helices

5) electrostatic attraction between charged amino acid side chains

**(1) Chptr 22**

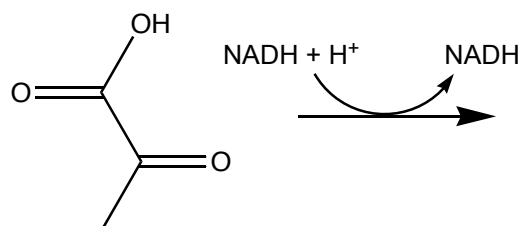


30. (2 points) Which structural element(s) can stabilize polar groups in the interior of a protein (choose the best answer)?

- 1) quaternary structure  
 2) disulfide bonds  
 3) alpha helices  
 4) beta sheets  
 5) alpha helices and beta sheets

**(5) Chptr 22**

31. (2 points) In one of the reactions of glycolysis, pyruvate reacts with NADH. What is the structure of the product?



- (1) (2) + (3) (4) +

**(1) NADH carries out reductions (but only one) (Chptr 28)**

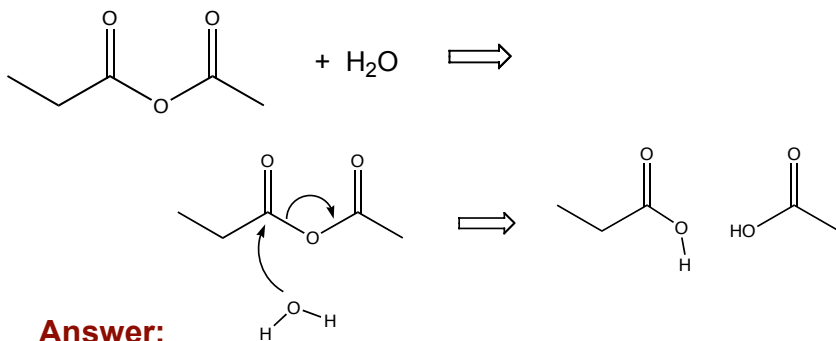
32. (2 points) What is the course number of this class?

- 1) 111                      2) 250                      3) 496                      4) 728

**(2)**

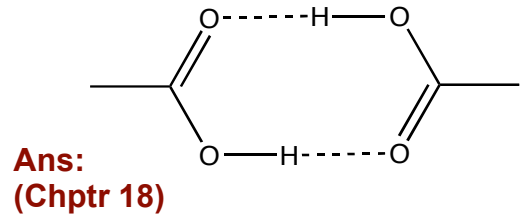
**\*\* Answer questions 32-39 directly on this sheet, in the spaces provided \*\***

33. (5 points) Draw the structural formula for the major organic product of the reaction below:

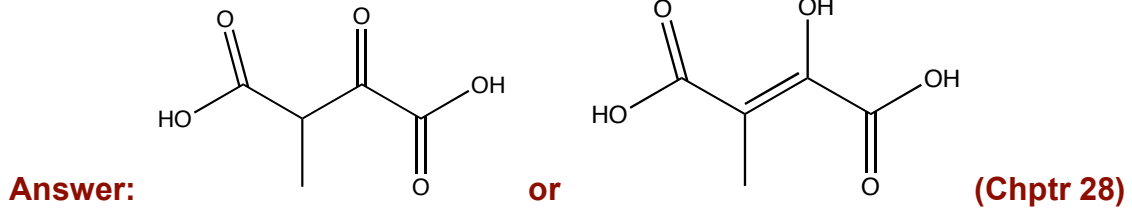
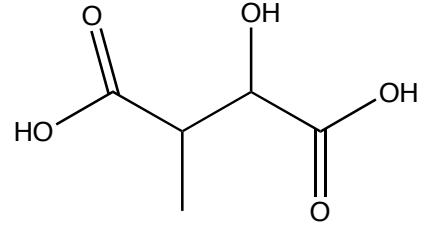


**OWL19.3b / Quiz 2**

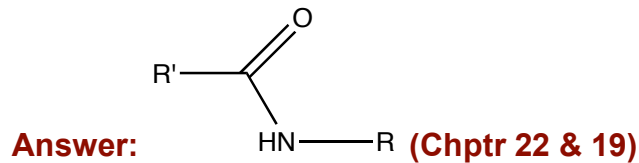
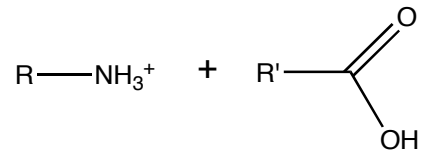
34. (5 points) Draw a structural formula for the dimer formed when two molecules of acetic acid (ethanoic acid) interact by hydrogen bonding.



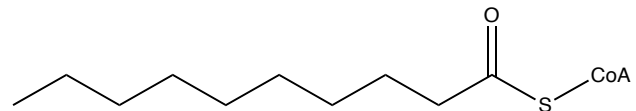
35. (5 points) In the Citric Acid cycle, which is a product of the reaction of malate (shown at right) with  $\text{NAD}^+$ ? (Hint #1: you are not expected to know this from memory, but deduce it from chemistry; Hint #2: there is no decarboxylation)



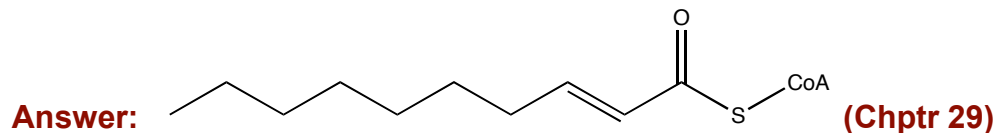
36. (5 points) Draw the product(s) of the following reaction:



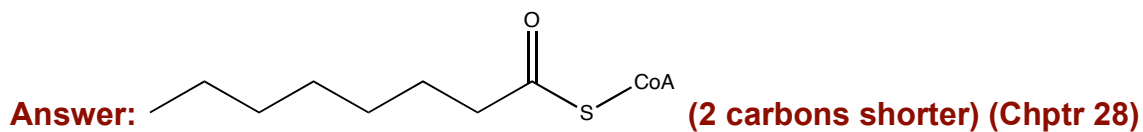
37. (5 points) For the active fatty acid at right,



Draw the structure of the immediate product of reaction with Acyl-CoA dehydrogenase.

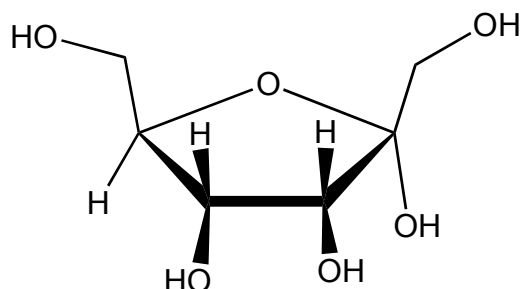


38. (5 points) For the same starting fatty acid above, draw the structure of the active fatty acid resulting from one complete round of  $\beta$ -oxidation



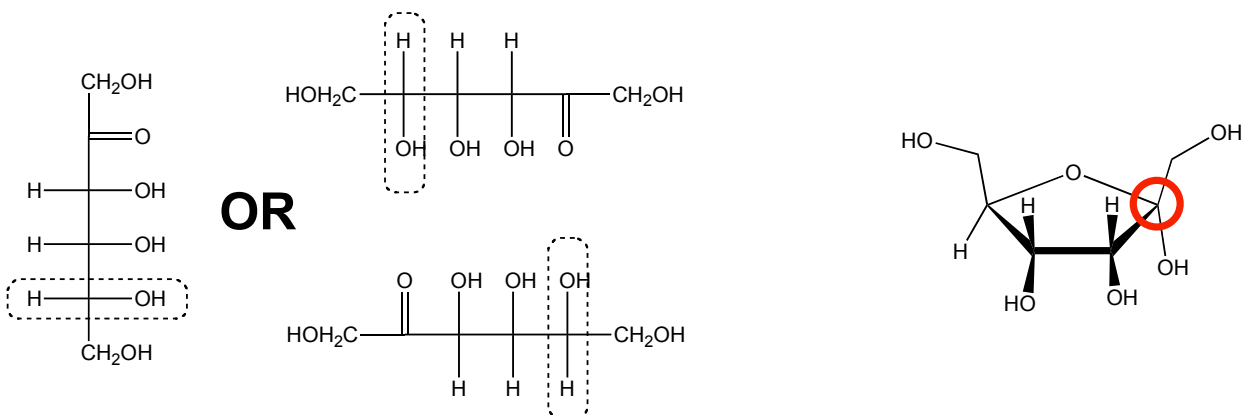
39. (3 points) Circle the anomeric carbon in the sugar at right.

**Chapter 20**



40. (5 points) Draw the Fischer projection corresponding to the linear form of the molecule shown at right.

**Chapter 20**



**Assigning the stereochemistry of the center in the dashed box is an advanced topic, and so is not graded on this exam.**