 Name: _____

Chem 111 *9:05a section* *Evening Exam #2v3*

This exam is composed of **25** questions. Go initially through the exam and answer the questions you can answer *quickly*. Then go back and try the ones that are more challenging to you and/or that require calculations.

As discussed on the course syllabus, honesty and integrity are absolute essentials for this class. In fairness to others, dishonest behavior will be dealt with to the full extent of University regulations.

$E = h\nu = \frac{hc}{\lambda}$

1 mL = 1 cm^3

Hz = s^{-1}

$h = 6.626x10^{-34}\ J\ s$

$c = 2.998x10^{8}\ m\ s^{-1}$

$N = 6.022x10^{23}\ mol^{-1}$

1. How many valence electrons are in the P atom?

   1) 3 2) 6 3) 5 4) 10 5) 0

2. Which atom(s) has/have completely filled 3s, 3p, and 3d orbitals?

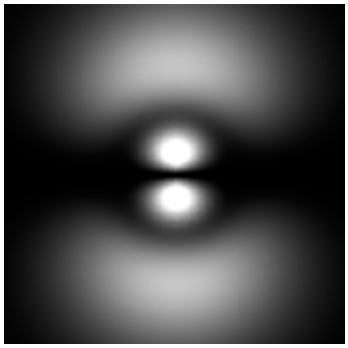
   1) Ar 2) Zn 3) Kr 4) Ar & Zn 5) Kr & Zn

3. Which element is represented by: $\mathbf{1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^4}$

   1) Sb 2) Te 3) Br 4) As 5) Se

4. The orbital depicted at right is:

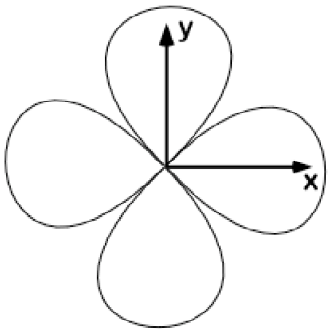
   1) 1s 2) 2p 3) 3s 4) 3p 5) 4p



 Name: _____

5. The orbital depicted at right is:

   1) p_{xy} 2) d_{xy} 3) $d_{x^2-y^2}$ 4) d_{z^2} 5) f_{xy}



6. Which of the following quantum number sets is ***not*** allowed?

   1) n=+3 l=+2 m_l = -1 m_s = +1/2 2) n=+2 l=+1 m_l = -1 m_s = +1/2

   3) n=+3 l=+1 m_l = -2 m_s = -1/2 4) n=+2 l=0 m_l = 0 m_s = +1/2

   5) n=+3 l=0 m_l = 0 m_s = -1/2

7. What is the maximum number of orbitals that can be identified by the set of quantum numbers n=+3 l=+1 ?

   1) 2 2) 7 3) 5 4) 6 5) 3

8. Which of the following has the shortest bond length?

   1)HF 2) H_2O 3) NH_3 4) CH_4

9. Which of the following has the highest bond energy?

   1) B_2 2) C_2 3) F_2 4) O_2 5) N_2

 Name: _____

10. The CO bond in the molecule CH_2O is best described as a:

    1) single bond
    2) double bond
    3) triple bond
    4) ionic bond
    5) the molecule doesn't exist

11. Consider the molecule SO_3^x, where x is the charge on the molecule. Two bonds are single bonds, one is a double bond. Which value of x yields the stable molecule? (Hint: draw Lewis structures to figure this one out)

    1) +2 2) +1 3) 0 4) –1 5) –2

12. For the SO_3^x molecule above, how many equal-energy resonance structures can you draw?

    1) 1 2) 2 3) 3 4) 4 5) 6

13. The NO bond in HNO is a:

    1) single bond 2) double bond 3) triple bond 4) ionic bond

14. If an element with the valence configuration $\mathbf{4s^2 3d^7}$ loses **2** electron(s), these electron(s) would be removed from the following **subshell(s)**.

    1) 4s 2) 3d 3) 4s and 3d 4) 3p 5) 4p

 Name: _____

15. Which molecule below does not exist?

1) BeF_2 2) CaF_2 3) Mg_2O 4) KCl 5) BCl_3

16. Draw a stable Lewis structure for the symmetrical molecule **hydrazine** N_2H_4. In this structure, how many ***lone pair electrons*** are on ***each*** N?

1) 0 2) 1 3) 2 4) 4 5) 6

17. Draw a stable Lewis structure for the molecule **OCS**. In this structure (with C at the center), what is the bond order between C and O?

1) 1 2) 2 3) 3 4) 0.5 5) 1.5

18. Draw the best Lewis structure for $\mathbf{ClF_4^-}$. How many **lone pair electrons** are located on Cl?

1) 1 2) 2 3) 3 4) 4 5) 6

19. For the molecule $\mathbf{ClF_4^-}$, what is the electron group geometry of Cl?

1) linear 2) tetrahedral 3) trigonal planar
4) trigonal bipyramidal 5) octahedral

 Name: _____

20. In the molecule $\mathbf{NO_2^+}$, the actual bond order for each NO bond is:

1) 1 2) 2 3) 3 4) 1.5

5) 1 for one bond and 2 for the other

21. Draw the Lewis structure for $\mathbf{XeF_4}$. The electron group geometry is:

1) octahedral 2) square pyramidal 3) trigonal bipyramidal

4) square planar 5) none of the above

Bond Dissociation Energies (kJ mol^{-1}) (gas phase)

Bond	D	Bond	D	Bond	D
H-H	436	C-C	346	N-N	163
C-H	413	C=C	610	N=N	418
N-H	391	O-O	146	C-O	358
O-H	463	O=O	498	C=O	745

22. Consider the reaction: H_2CCH_2 (g) + H_2 (g) → CH_3CH_3 (g)

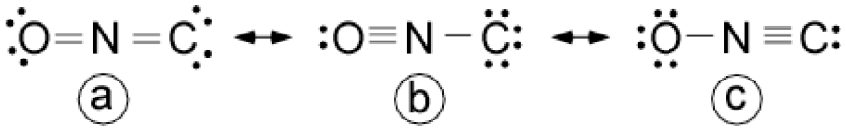
What is the energy ($\Delta H°$, in kJ mol^{-1}) for this reaction?

1) –346 2) +346 3) –44 4) +44 5) –480

5) –1 for one O and 0 for the other O

 Name: _____

(Questions 23-24) Consider the following resonance forms for the ion ONC^-



23. In resonance structure **b**, what is the formal charge on C?

1) –3 2) –2 3) –1 4) 0 5) +1

24. Which resonance structure is lower in energy, **b** or **c**?

1) b 2) c 3) neither, they have the same energy

25. The correct designator for this course is:

1) Chem 3.14 2) Chem 363 3) Chem 111 4) Sports 01

PERIODIC TABLE OF THE ELEMENTS

1A	2A	3B	4B	5B	6B	7B	8B	8B	8B	1B	2B	3A	4A	5A	6A	7A	8A
1 H 1.008																	2 He 4.003
3 Li 6.939	4 Be 9.012											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 19.00	10 Ne 20.18
11 Na 22.99	12 Mg 24.31											13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.07	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.90	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.71	29 Cu 63.55	30 Zn 65.39	31 Ga 69.72	32 Ge 72.61	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc (99)	44 Ru 101.1	45 Rh 102.9	46 Pd 106.4	47 Ag 107.9	48 Cd 112.4	49 In 114.8	50 Sn 118.7	51 Sb 121.8	52 Te 127.6	53 I 126.9	54 Xe 131.3
55 Cs 132.9	56 Ba 137.3	57 La 138.9	72 Hf 178.5	73 Ta 181.0	74 W 183.8	75 Re 186.2	76 Os 190.2	77 Ir 192.2	78 Pt 195.1	79 Au 197.0	80 Hg 200.6	81 Tl 204.4	82 Pb 207.2	83 Bi 209.0	84 Po (209)	85 At (210)	86 Rn (222)
87 Fr (223)	88 Ra 226.0	89 Ac 227.0	104 Unq (261)	105 Unp (262)	106 Unh (263)	107 Uns (262)	108 Uno (265)	109 Une (266)									